

# Ignition of Wood: A Review of the State of the Art

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**ABSTRACT:** This review encompasses the available practical and experimental data on the ignition of solid wood. Only solid, natural wood is considered, not sawdust, chips, or products that have been treated with fire retardants or other substances, nor the ignition of living trees. Panel products such as plywood or particleboard have ignition properties very similar to solid wood, so the solid-wood results will generally be applicable to them. Wood may ignite by flaming directly, or it may ignite in a glowing mode, which may or may not be followed by flaming. It is shown that the ignition temperature is around 250°C for wood exposed to the minimum heat flux possible for ignition, and that it invariably ignites, at least initially, in a glowing mode under these conditions. The ignition temperature rises rapidly as the heat flux is increased. Piloted ignition at heat fluxes sufficient to cause a direct-flaming ignition normally occurs at surface temperatures of 300–365°C. Autoignition temperatures at fluxes higher than minimum are essentially unknown. No theory is available that encompasses the possibility of glowing, glowing followed by flaming, or direct-flaming ignition modes. Most published studies have dealt with radiant or radiant + convective heating, and knowledge is extremely poor for ignition from direct contact by hot bodies or by flames. A species-independent correlation is derived for the radiant, piloted ignition of thermally-thick wood, but the fit is only fair. The minimum flux for ignition is  $4.3 \text{ kW m}^{-2}$ , based on a single study; most reported tests have been much too brief to produce useful data on this point.

**KEY WORDS:** Cone Calorimeter, glowing combustion, ignition, Setchkin furnace, wood.

## IGNITION TEMPERATURE

**T**HE CONCEPT THAT combustible substances ignite when a given surface temperature is first attained is an empirical notion – in many cases, this is found to be true enough, so that even though not exactly true, the concept

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has utility and merit. It has also found significant application to theoretical modeling – closed-form theories for radiant ignition, for example, generally assume that ignition corresponds to a known, constant surface temperature  $T_{ig}$ . However, it is often considered that there are actually two different  $T_{ig}$  values, one for autoignition (unpiloted) and another for piloted ignition. Thus, the starting point for investigating the ignition of wood must be to examine experimental data on its ignition temperature. As can be seen in Table 1, studies on this question go back well into the 19th century and have continued until the present time. The spread of data is clearly enormous. It might first be noted that even the term ‘ignition temperature’ tends to mean two different things: (1) the temperature of the surface at the time of ignition; or (2) the minimum temperature of a furnace sufficient for a specimen put therein to ignite. The latter notion might seem to be old and non-rigorous, but it must be remembered that: (a) the common test for ignition temperature is the Setchkin furnace, ASTM D 1929 [1], which is based on the latter definition; and (b) the user often needs to know the highest environment temperature to which he can subject a material without it igniting and he may be less interested in actual temperatures at the specimen. Excluding one value, the results in Table 1 span 210–497°C for piloted ignition and 200–510°C for autoignition. The following reasons should be considered that might account for the spread:

- the definition of ignition that is used
- piloted versus autoignition conditions
- the design of the test apparatus and its operating conditions
- specimen conditions (e.g., size, moisture, orientation)
- species of wood.

The definition of ignition is complicated not only by the two meanings currently in use, but also by some practices followed by earlier investigators. Until the 1960s or so, it was not rare for investigators to report ignition results without making visual observations. Strange as this may seem from today’s perspective, a number of studies exist where the ignition criterion was based solely on thermocouple readings. Typically, the test rig was equipped with two thermocouples and a criterion was used which related the value or slope of the one reading to the other. Results of this kind might be automatically excluded from consideration, except for the fact that data from those investigators does not seem to be systematically different from the others.

The design of the test apparatus may have a significant influence. The majority of devices fall into one of two types: (1) a furnace into which a small specimen is bodily plunged; or (2) a specimen sitting in the open air and being radiatively heated, e.g., the Cone Calorimeter [34]. But this basic

**Table 1. Summary of ignition temperature results for wood.**

Year	Investigator	Spec. size	Ignition temperature (°C)		Comments
			Piloted	Auto-ignition	
1887	Hill [2]	0.5–15 g		<b>220–300</b>	measured air temperature near sample
1910	Bixel, Moore [3]	35 mm ?		200–250	measured oven temperature; scant details
1922	Banfield, Peck [4]	50 × 50 × 200 mm		302–308	measured surface temperature
1934	Brown [5]	1–5 g		<b>220–250</b>	measured oven temperature; tiny samples; unsound ignition criterion
1936	VanKleeck [6]	chips		<b>235</b>	measured specimen temperature; unsound ignition criterion
1947	NIST [7]	shavings		<b>228–264</b>	softwood shavings in test tube; criterion – glowing or flaming
1949	Graf [8]	7–13 g		<b>232–245</b>	measured oven temperature; tiny samples; unclear ignition criterion
1949	Angell et al.[9]	13 × 19 × 51 mm		204	measured gas temperature close to specimen
1950	Fons [10]	2–9 mm cylinders		343	measured oven temperature; solved inverse problem
1958	Narayanamurti [11]	?	228		measured oven temperature
1959	Thomas et al. (data of Prince, 1915) [12]	32 × 32 × 102 mm	210		measured oven temperature; solved inverse problem
1959	Akita [13]	20 × 20 × 1.8 mm	450 < 350	489	measured oven temperature; solved inverse problem
1960	Simms [14]	8 mm $\phi$		545	measured oven temperature only; calculated from correlation, not measured
1960	Moran [15]	50 × 50 × 6.4 mm		<u>255</u>	at flux = 25 kW m <sup>-2</sup> ; measured surface temperature
1961	Patten [16]	3 g shavings	<b>260</b>	<b>260</b>	measured oven temperature (Setchkin test)

(continued)

Table 1. Continued.

Year	Investigator	Spec. size	Ignition temperature (°C)		Comments
			Piloted	Autoignition	
1961	Buschman [17]	57 × 57 × 8 mm	<u>369</u>		calculated from correlation; fluxes 14.3–37.2 kW m <sup>-2</sup>
1964	Shoub, Bender [18]	920 × 920 mm		<u>254</u>	measured surface temperature
1964	Tinney [19]	≥ 6 mm $\phi$		<u>350</u>	measured oven temperature
1967	Simms, Law [20]	76 × 76 × 19 mm	<u>380</u>		calculated from correlation
1967	Muir [21]	80 × 160 mm	<u>364–384</u>		measured surface temperature; flux range 15–25 kW m <sup>-2</sup>
1967	Koohyar [22]	100 × 100 × 12–19 mm	<u>361</u>	<u>402</u>	measured surface temperature; flux range 18–35 kW m <sup>-2</sup> (piloted), 32–35 kW m <sup>-2</sup> (autoignition)
1969	Melinek [23]	100 × 100 × 13 mm	<u>353</u>	<u>382</u>	calculated from correlation
1969	Jach [24]	few grams		<b>260–290</b>	measured oven temperature
1970	Smith and King [25]	75 × 75 × 19 mm	<u>350</u>	<u>413–714</u>	temperatures measured by optical pyrometry; autoignition values dubious
1983	Atreya [26]	64 mm $\phi$ × 19 mm	<u>370</u>		temperatures measured, but below surface; flux = 18 kW m <sup>-2</sup>
			<u>350</u>		temperatures measured, but below surface; flux ≥ 30 kW m <sup>-2</sup>
1986	Atreya et al. [27]	75 × 75 × 19 mm	<u>330–405</u>		temp. measured, but below surface

1988	Abu-Zaid [28]	150 × 75 × 37 mm	<u>420</u>	forced-air flow; temp. measured but below surface; flux = 18.5 kW m <sup>-2</sup>
			<u>350</u>	forced-air flow; temp. measured but below surface; flux > 25 kW m <sup>-2</sup>
			<u>530</u>	flux = 40 kW m <sup>-2</sup>
1991	Janssens [29]	100 × 100 × 17 mm	<u>300–364</u>	surface temp. measured; fluxes 25–35 kW m <sup>-2</sup>
1992	Li, Drysdale [30]	64 × 64 × 18 mm	<u>411–497</u>	temp. measured but below surface; flux < 20 kW m <sup>-2</sup>
			<u>353–397</u>	temp. measured but below surface; flux > 20 kW m <sup>-2</sup>
1993	Masařík [31]	2.5 g	<b>220–240</b>	tested wood fiberboard; measured oven temperature (Setchkin test)
1996	Fangrat et al. [32]	100 × 100 mm	<u>296–330</u>	surface temp. measured; fluxes ≥ 25 kW m <sup>-2</sup>
1997	Moghtaderi et al. [33]	100 × 100 × 19 mm	<u>332</u>	temp. measured but below surface; at 20 kW m <sup>-2</sup>
			<u>297</u>	temp. measured but below surface; at 60 kW m <sup>-2</sup>

? = unknown measurements; **Bold** = Type 1 values (see text below); Underline = Type 2 values (see text below).

**Table 2. Summary of ignition temperature data.**

Type of Test	Ignition Temperature (°C)	
	Piloted	Autoignition
A few grams plunged into a furnace	220–260	220–300
Radiant heating of a largish specimen	296–497	254–530
others; unidentified	210–450	200–525

division is confounded by the fact that there is a preferred specimen type for each test: specimens of only a few grams are normally put into a furnace that exposes the whole specimen bodily, while specimens placed in front of radiant heaters are typically on the order of 100 g and of sizeable dimensions in at least two directions. The results are summarized in Table 2, with Type 1 values indicated in **bold** in Table 1 and Type 2 underlined.

Considering first autoignition temperatures under radiant heating, the results evidently span a huge range. Smith's [25] results (which go up to 714°C) appear to be implausible and may refer to an average optically measured temperature on which some spots are already glowing; thus, they will be excluded. Several other workers reported calculated, rather than measured, values; these will be presumed to be less reliable. Of the values listed in Table 1 that are measured (rather than imputed), Moran's [15] value of 255°C and Shoub's [18] 254°C are impressively close. The only other values obtained by actual measurement are from Koohyar [22] and Abu-Zaid [28]. Koohyar [22] obtained an average value of 402°C over the flux range of 32–35 kW m<sup>-2</sup>, while Abu-Zaid's [28] determined 530°C at a heat flux of 40 kW m<sup>-2</sup>. Since both of the latter measurements were obtained at fluxes higher than Shoub's [18] 4.3 kW m<sup>-2</sup> or Moran's [15] 25 kW m<sup>-2</sup>, this suggests that separate flux regimes must be considered. Thus, it might be assumed that 250°C is characteristic at very low fluxes, while some much higher temperature is obtained at high heat fluxes. Turning now to autoignition in 'a few grams plunged into a furnace' tests, if the range reported by each investigator is averaged, then the data spans only 235–275°C, with an average of exactly 250°C. It may be noted that the 'a few grams plunged into a furnace' tests are normally operated in such a way as to only seek out the condition where the furnace temperature is the minimum for ignition. In principle, they can be run at nonminimum temperatures, but such data are hardly ever reported. Thus, from this type of test, there is no corresponding result to the high-flux region of radiant tests. It can be concluded then that if a wood specimen is ignited under external heating barely sufficient to ignite it, it will ignite at approximately 250°C regardless of the type of heating arrangement.

Concerning autoignition at higher heat fluxes, the paucity of data makes it difficult to draw firm conclusions. Koohyar's [22] value of 402°C is probably reliable. Simms' [20] calculated value of 545°C is close to Abu-Zaid's [28] measured 530°C, but both seem very high. Koohyar [22] has pointed out that Simms [20] made a mistake in his calculations and that the correctly-deduced value would have been around 485°C. Akita's [13] calculated value of 489°C also appears to be too high (see below). In general, it is evident that apparatus details play a stronger role in autoignition than for piloted ignition, leading to wider scatter.

For piloted ignition,  $T_{ig}$  values should not be any higher than those for autoignition. The only way that the converse could be true is either due to natural data scatter, or if the equipment is so badly designed that the pilot actually interferes with ignition. Only two workers have presented 'a few grams plunged into a furnace' data for piloted  $T_{ig}$ . The values are 260°C from Patten [16] and 220–240°C from Masařík [32], giving an average of 245°C, which can be taken as identical to 250°C. The conclusion is that piloting does not make any difference on  $T_{ig}$  in tests of this type. Considering next piloted ignition results from radiant heating tests, it is evident that none are available at heating conditions barely enough for ignition. The available results are typically for specimens 12–25 mm thick and exposed for only 10–60 min. Shoub's [18] data indicates that much longer times are needed for specimens of these thicknesses before minimum conditions are approached. On the basis that piloted values should not be lower than autoignition,  $T_{ig} = 250^\circ\text{C}$  can be provisionally assigned also as the piloted ignition temperature for radiant tests. Thus, it is concluded that 250°C is the best estimate of the ignition temperature irrespective of piloting and irrespective of type of test, provided that heating conditions are just barely enough for ignition.

At this point, it is important to observe the nature of the low-heat ignitions. Moran [15], Li [30], and Spearpoint [35] all describe the same phenomenon: ignition starts as a glowing ignition\* and flaming is seen later, if at all. By the way, the glowing ignition temperature must not be confused with the temperature of the glowing spot. In a glowing ignition, a glow begins at one spot and very quickly reaches red-hot conditions (over 600°C). This high temperature is not the glowing ignition  $T_{ig}$ ; instead the latter must be determined either by a thermocouple reading just before a steep jump takes place or by a thermocouple on the same surface but away from the spot of initial glow. The glowing ignition phenomenon also serves to explain why no difference is seen between autoignition and piloted ignition results. If flaming

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\*Glowing combustion entails incandescence without flaming. It is a direct surface oxidation process, but differs from smoldering in that smoldering is a self-sustained process whereas glowing combustion is supported by external heating.

is preceded by glowing, then the glowing zone can serve as a high temperature pilot, if subsequently sufficient pyrolysates emerge to be ignitable as a flame. We shall refer to the regime where glowing occurs first, followed by flaming, as *2-step ignitions*. Parenthetically, unlike wood, materials that are not susceptible to glowing ignition (e.g., thermoplastics) show a substantially lower  $T_{ig}$  in the Setchkin furnace for piloted than for autoignition conditions.

'A few grams plunged into a furnace' tests generally share two features: very small specimen size, and exposure to conditions where, apart from radiant heating, the specimen is convectively heated (by contrast, in radiant heating tests the convective stream is cooling the specimen). There is one test series where fairly sizeable specimens were plunged into a furnace, and that is Prince's 1915 study [36]. His original study reported that ignition was attained for furnace temperatures of 180–200°C. Thomas [12] later estimated specimen surface temperatures by modeling and concluded that surface temperatures at ignition were 30°C higher than the furnace temperature. This correction, which arises due to self-heating, is negligible for tiny specimens and increases with increasing specimen size. Since Thomas' [12] corrected values are in the same range as the raw values from 'few grams' specimen tests, the conclusion is that there is no specimen size dependence, at least when testing under heating conditions barely sufficient to cause ignition.

Considering next piloted ignition at higher fluxes, only the radiant tests can be considered, since 'a few grams plunged into a furnace' tests are not normally run this way. For such heat fluxes,  $T_{ig} \approx 300\text{--}350^\circ\text{C}$  covers all, or nearly all results of Janssens [29], Atreya [26], Abu-Zaid [28], Fangrat [32] and Moghtaderi [33]. Akita's [13] value of 450°C, obtained by calculation, appears to be wrong since he did obtain ignitions at a furnace temperature of 350°C (and did not try lower temperatures). For his 1.8 mm thick specimens, self-heating would be minimal, so the actual ignition temperature appears to have been below 350°C, making his measurements also consistent. Of modern workers with good equipment, only the results of Li and Drysdale [30] are outside this range and these are about 50°C higher, for unknown reasons. Janssens [29] noted that the range can be further shrunk by considering the slight but systematic effect of wood type. His results for oven-dried specimens were: hardwoods 300–311°C; softwoods 349–364°C. At fluxes high enough to ensure a direct-flaming ignition, these values can be adopted for piloted  $T_{ig}$ . Wood is comprised of three primary constituents – cellulose, hemicellulose, and lignin. Hemicellulose ignites at the lowest temperature, lignin higher, and cellulose higher yet [37]. Compared to hardwoods, softwoods have a smaller fraction of hemicellulose and a higher fraction of lignin, thus accounting for their higher  $T_{ig}$ .

Next the intermediate-flux regime must be considered, where the heat flux is higher than the minimum flux, but is low enough for ignitions to be of the



2-step type. This reported data spans a sizeable range of 332–497°C. Part of the scatter is probably due to experimental difficulties, since Urbas and Parker [38] observed that considerable care needs to be exercised to properly instrument a surface that is undergoing charring. Part of the difference, however, is real and is attributable to changed exposure conditions. Moran's [15] data are instructive here. Although intermediate data were scattered, as the flux was raised from 25 to 29 kW m<sup>-2</sup>, the ignition temperature rose from 255 to 301°C while the ignition time dropped by 33%. The reason for the dependence of  $T_{ig}$  on flux in this regime will be considered in the next section. The 300°C value is significant, since wood pyrolysis involves competing mechanisms, with temperatures under 300°C leading largely to charring, while over 300°C gasification being favored [39]. Thus, if heating conditions are such that the material does not exceed 300°C, a glowing ignition is favored.

Not much data is available to suggest the flux level (or if, indeed, there is a fixed flux value) which separates 2-step from direct-flaming ignitions. Koohyar's [22] data on piloted ignition shows 2-step ignitions over the flux range of 18–31 kW m<sup>-2</sup>, but his lowest direct-flaming ignition was at 20 kW m<sup>-2</sup>. His autoignition data shows 2-step ignitions over 32–35 kW m<sup>-2</sup>, with direct-flaming at a flux as low as 31 kW m<sup>-2</sup>. The pilot ignition data of Li and Drysdale [30] were taken only at coarse flux intervals, but indicate 2-step ignitions below 20 kW m<sup>-2</sup> and direct-flaming at 24 kW m<sup>-2</sup> and higher. If a cross-draft is present, the flux range over which 2-step ignitions occur is expanded [40].

Concerning other systematic effects, at the minimum flux condition, Moran [15] found no difference in  $T_{ig}$  between oven-dried and room conditioned specimens. In the medium flux regime under piloted conditions, Janssens [41] concluded that  $T_{ig}$  rises by 2°C for each percent of moisture content increase. This will normally be insignificant for practical moisture contents. Grain orientation (i.e., along-grain vs. end-grain exposure) may also have an effect on  $T_{ig}$ , but good enough data are not available to explore the issue. Almost all existing experimental data deal with along-grain exposures, which are also common in accidental fires.

## GLOWING IGNITION MODELING

A glowing ignition involves the direct surface oxidation of a material (heterogeneous reaction), thus Baer and Ryan [42] suggested that the simplest model for this is:

$$\frac{\partial T}{\partial t} = \frac{\lambda}{\rho C} \frac{\partial^2 T}{\partial x^2}$$

with the boundary condition:

$$-\lambda \left. \frac{\partial T}{\partial x} \right|_{x=0} = \dot{q}_e'' + B_s Q_s \exp(-E_s/RT)$$

where  $T$  is the temperature;  $t$ , time;  $\lambda$ , thermal conductivity;  $\rho$ , density;  $C$ , heat capacity;  $\dot{q}_e''$ , irradiance;  $B_s$ , preexponential factor;  $Q_s$ , heat of reaction;  $E_s$ , activation energy for surface reaction; and  $R$ , universal gas constant. Based on this, Lengellé et al. [43] then showed that a solution for the ignition temperature  $T_{ig}$  is:

$$T_{ig} = \frac{E_s}{R} \left[ \ln \left( \frac{B_s Q_s}{\alpha \dot{q}_e''} \right) \right]^{-1}$$

where  $\alpha$  = nondimensional temperature rise associated with ignition. The  $\alpha$  factor serves as an ignition criterion and they found empirically that  $\alpha \approx 0.15$  corresponds to ignition. The equation shows that  $T_{ig}$  decreases with decreasing irradiance, and Lengellé [43] demonstrated that this indeed occurs experimentally for a number of propellants. Propellants are, of course, substances very different from wood, but Moussa et al. [44] proposed that the same equation be used in describing char oxidation occurring during smoldering of wood; however, they did not provide quantitative values for the kinetic constants. Fredlund [45] used a slightly different term in his model of wood combustion, but provided no experimental verification in the glowing ignition regime. Bilbao et al. [40] described an elaborate model with multiple, piecewise linearized kinetics regimes, but the model's ability to predict glowing ignition temperatures was not reported. Ohlemiller [46] noted that describing char oxidation of wood is difficult, since the char is not a unique chemical entity, but rather, is a substance whose characteristics are history-dependent. For a similar material, coal char [47], the chemical properties are, in fact, strongly dependent on the physical nature (pore structure) of the char that has been created, and it might be expected that this would also be important for wood. More complex heterogeneous reaction models that include pre-structure effects (and the possibility of both kinetically-limited and diffusion-limited reaction rates within these structures) are available for coal-char combustion [48], but such models have yet to be applied towards representing the ignition of wood. The above observations help to place in context the long times required for glowing ignition of wood – plywood required over 5 h in Shoub and Bender's [18] experiment. This long time period is associated with creating of a reactive porous char. The conclusion,

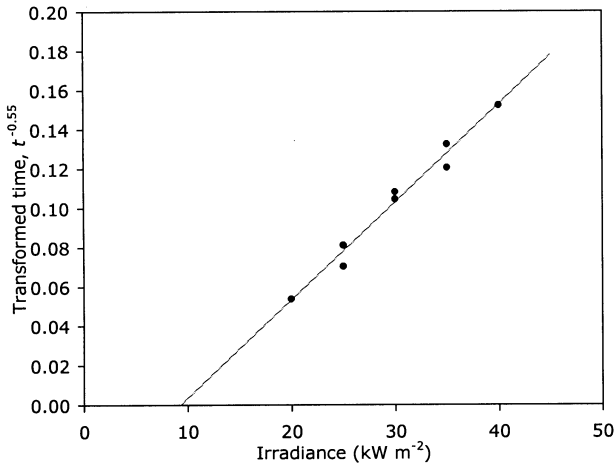
thus, has to be that only qualitative rudiments are known for glowing ignition, and that quantitative modeling is not yet possible, largely because of an absence of experimental data.

### FLAMING IGNITION FROM RADIANT HEATING

#### Theory

There is more than half a century of history in the development of both comprehensive and ‘practical’ theories of flaming ignition of wood materials. Janssens [29,41] reviewed them extensively and here only the salient feature will be reprised: his recommended method for plotting experimental data so that sound interpolations and extrapolations may be possible. His study, which was based on numeric approximations to an inert-body model of an igniting solid, entails plotting the ignition time raised to the  $-0.55$  power on the  $y$ -axis and the external imposed heat flux (irradiance) on the  $x$ -axis. This is illustrated with Janssens’ own data in Figure 1. Since a straight line can be obtained when the data is plotted in this way, only two parameters are needed to describe the data fit. An obvious one to choose is the  $x$ -axis intercept, denoted as  $\dot{q}''_{cr}$ . The slope is a very small number, so it is more convenient to select the inverse of the slope and to designate it as  $B_{ig}$ . Thus, the equation describing the data plot is:

$$t_{ig}^{-0.55} = [\dot{q}''_e - \dot{q}''_{cr}] / B_{ig}$$



**Figure 1.** Janssens’ piloted ignition results for *Blackbutt*, oven-dried, vertical orientation.

In the example,  $\dot{q}_{cr}'' = 9.3 \text{ kW m}^{-2}$ ,  $B_{ig} = 201 \text{ kW m}^{-2} \text{ s}^{+0.55}$ . In general, it is found that ignition may not be possible at fluxes just slightly greater than  $\dot{q}_{cr}''$ , and a higher heat flux is necessary for ignition to actually occur. This latter value is designated  $\dot{q}_{min}''$ , the minimum flux for ignition. Thus, apart from the two parameters needed to describe the straight line, a third parameter is needed which denotes the lowest point on the line that has physical meaningfulness. Janssens [29,41] presented a second method for thermally thin materials. Physically, whole wood is rarely used as free-standing in minuscule thicknesses (e.g.,  $< 1 \text{ mm}$ ), thus Janssens' second procedure will not be presented here. But the 'thermal thickness' is not necessarily the same as the physical thickness, and substances of finite thickness which behave as thermally-thick bodies when initially heated will eventually respond as thermally-thin, if sufficient time has elapsed. This point is treated in the next section.

Janssens' theory was mainly intended as an aid to using experimental data and was not intended to encompass all relevant physicochemical phenomena. Indeed, since it is an inert-solid theory, events in the gas phase are ignored and ignition is assumed to uniquely occur at the moment a certain face temperature is first attained. Much more refined theories have been put forth in recent years, for example, Yuen's [49]. These have the limitations that they (a) require a large amount of input data, much of which may be unavailable or uncertain; and (b) difficult numeric computations must be performed for each problem; consequently, they are not useful as 'data plotting aids.' While advanced theories attempt to capture gas-phase ignition events, there is currently no theory available, simple or complex, which encompasses the possibility that a specimen may exhibit glowing ignition, glowing  $\rightarrow$  flaming (2-step) ignition, or a direct-flaming ignition.

### Experimental Results on Piloted Ignition

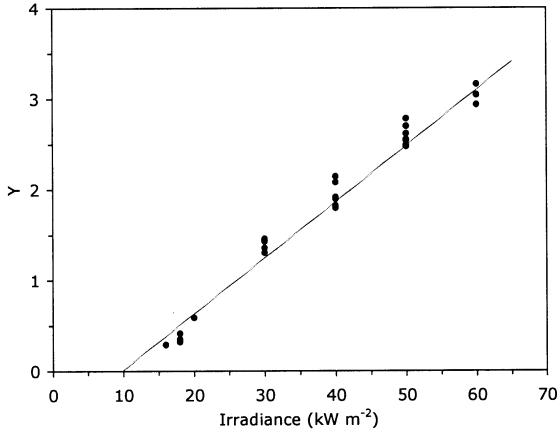
From both theory and experiments, it is evident that a number of variables can affect the ignition time of thermally-thick, solid wood, of which density, thermal conductivity, moisture content, and geometric factors are probably the most important. Taking the last first, in testing, geometric effects show up as apparatus-dependent factors, since no physical test rig can capture apparatus-independent properties of a material. Size of specimen is a geometric variable to consider, but Long et al. [50] noted that the only scale-dependent term in basic ignition theory is the convective heat transfer coefficient,  $h_c$ , which varies with size  $L$  according to  $h_c \sim L^{-1/4}$ . The effect on ignition time is much smaller than the change in  $h_c$ , since heat losses are dominated by radiation, and would be negligible for all except huge changes of scale. But basic ignition theory does not deal with events in

the gas phase and these may also have an effect. Within a single test apparatus, experimental data suggests that the size effect is very small [51], although when comparisons are made where both the scale and the basic apparatus are changed, somewhat larger differences crop up [52,53].

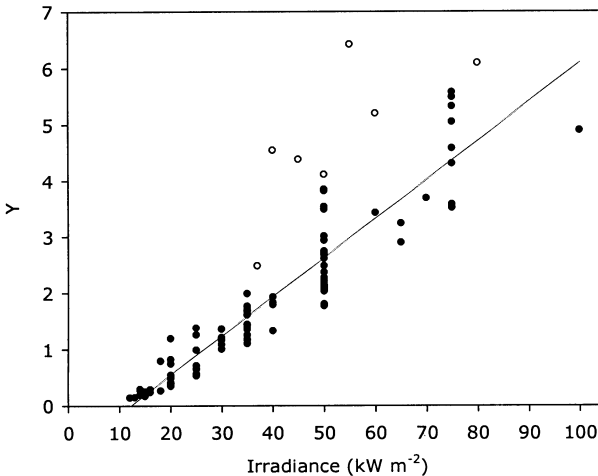
In any case, currently there is a sizeable database of test results only from the Cone Calorimeter, so for consistency, only Cone Calorimeter data obtained on samples exposed in the along-grain orientation will be considered. According to basic theory [68], for thermally-thick materials  $t_{ig} \sim \lambda \rho C (T_{ig} - T_o)^2$ . Since radiant ignition data are easy to obtain, but become much more difficult if surface temperatures need to be accurately measured, the consequence is that most investigators record only the flux and the ignition time. Consequently, it is best to treat  $T_{ig}$  from such data sets as part of the unknown constants to be fitted, thus taking  $t_{ig} \sim \lambda \rho C$ . Now, values of heat capacity,  $C$ , tend to vary little among members of a chemical family, and this appears to also be a reasonable conclusion for woods. Density, however, can vary by about a factor of 10, if exotic woods are included. Thermal conductivity increases with increasing density, with the simplest assumption being that  $\lambda \sim \rho^n$ , where the value of  $n$  remains to be determined. Thus, it seems appropriate to seek a correlation where  $t_{ig} \sim \rho^m$ , where  $m$  is also to be determined. This is not a novel idea, and Hallman [54] took a similar approach some 30 years ago.

Moisture content (MC) can have a complex effect, both because it directly affects the thermophysical properties and because, if it were to be treated accurately, an inert-substance model is no longer a viable starting point for a theoretical treatment. To make an accurate treatment of moisture, the extremes of green wood to oven-dried wood would have to be considered. Green wood can have  $MC > 100\%$ , but there are no available ignition data on it, with the literature containing data only for oven-dried specimens and ones that are equilibrated to room conditions. For room-conditioned wood specimens, MC depends on the humidity present, but across the U.S. it normally spans only the range of 4–14% [55], which is a small range and only covers the ‘zero-end’ of the scale. Most test results available are either for the oven-dried condition or for 9–12% moisture content, obtained by room-conditioning the specimens. As indicated above, Janssens concluded that moisture slightly increases  $T_{ig}$ , but this can be ignored unless the wood is green (for which no data are available, anyway).

To find a correlation, a large number of published [33,35,56–59] and unpublished [60–62] data sets were collected. These covered four test conditions: oven-dried horizontal, oven-dried vertical, room-conditioned horizontal, and room-conditioned vertical. Figure 2 shows the results for oven-dried horizontal specimens [59,61]; with one data set [33] not used due to excessive outliers. The densities spanned 170–850 kg m<sup>-3</sup>. Since this data



**Figure 2.** Correlation for oven-dried horizontal specimens.



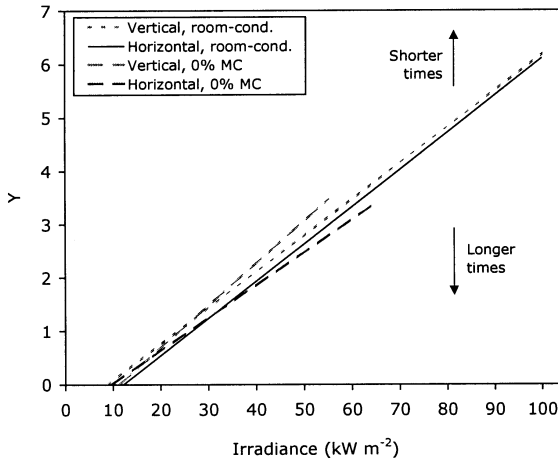
**Figure 3.** Correlation for room-conditioned horizontal data; hollow points were not used to derive the correlation.

set showed a relatively tight correlation, the exponent for the density term was derived from the data fit on this data set and fixed at that value for the remaining data fits. The value plotted on the  $y$ -axis is  $Y$ , which was taken as  $Y = t_{ig}^{-0.55} / \rho^{-0.4}$ . The other data sets showed higher scatter, for example, Figure 3.

Table 3 gives a summary of the correlations obtained. Figure 4 shows that the correlations are very similar and that it is reasonable to assign an ‘overall’ correlation. Clearly, a dry specimen ignites quicker than a moist

**Table 3. Summary of data correlations for piloted radiant ignition in the cone calorimeter.**

Conditions	$\dot{q}_{cr}''$ (kW m <sup>-2</sup> )	Constant	Tot. Data Points	Data Points Used
horiz., 0% MC	9.8	159	31	26
horiz., room	12.2	128	103	94
vert., 0% MC	11.5	99	67	48
vert., room	9.0	133	53	48
overall	11.0	130		



**Figure 4. Data correlations obtained for the four data sets.**

one, but this is somewhat violated in the correlations, and this is one reason why it is best to assign a single correlation, with the realization that moisture effects are swamped by general data scatter. It is also known that vertically-oriented specimens take longer to ignite than do horizontally-oriented ones [57] but, again, the scatter of the data does not permit this to emerge from the correlations. Based on these considerations, the estimating rule for radiant heating ignition of wood becomes:

$$t_{ig} = \frac{130\rho^{0.73}}{(\dot{q}_e'' - 11.0)^{1.82}}$$

According to theory, it would appear that the exponent for  $\rho$  is unusually low, but the reason for this is not clear. The root-mean-square error of the predictions is 64%, which indicates that predicting times to ignition can only

be done semi-quantitatively, but this must also be placed in the context that experimental data went from 2.5 to 4200 s, or a range of 1:1680. A close inspection of Figure 3 also reveals that below about  $15 \text{ kW m}^{-2}$ , the points deviate systematically above the straight line. This is as might be expected, since the theory is based on a thermally thick material, and wood specimens 12–25 mm thick no longer behave in a thermally thick manner when heated for a long time. It is possible to eliminate this systematic bias by fitting exponents higher than 1.82 to the irradiance factor, for example, 2.8 as suggested by Wesson [63]. But overall scatter still remains large and the treatment becomes wholly empirical.

The minimum flux for ignition is often the quantity of interest. In 1965, McGuire [64] suggested that this value can be taken as approximately  $12.5 \text{ kW m}^{-2}$  for most wood materials apart from low-density fiberboard. A value of  $12.5 \text{ kW m}^{-2}$  has subsequently been used for design purposes in many countries. This is indeed the value that is customarily obtained in the Cone Calorimeter and in other test methods where the time allotted for observation of ignition is 10–20 min. But lower values have been found, although not widely publicized. Spearpoint [35] recently explored both low-flux ignition and end-grain ignition of woods. Almost all ignition results available for wood are performed on specimens oriented towards the heat source along the grain, but different results are obtained when exposed to the end-grain. For along-grain exposures, Spearpoint found  $\dot{q}''_{\min} = 12.5 \text{ kW m}^{-2}$  for redwood and somewhat less than  $12 \text{ kW m}^{-2}$  for maple. But for end-grain ignition of maple, the lowest flux at which ignition occurred was  $8 \text{ kW m}^{-2}$ , with no ignition at  $7 \text{ kW m}^{-2}$ , making  $\dot{q}''_{\min} = 7.5 \text{ kW m}^{-2}$ . The minimum flux for end-grain ignition of redwood was not fully explored, but was found to be below  $9 \text{ kW m}^{-2}$ . For ignitions occurring at fluxes below  $10 \text{ kW m}^{-2}$ , a glowing ignition preceded flaming. The times associated with the low-flux ignitions were notably long, it taking 2680 s for end-grain ignition of maple at  $8 \text{ kW m}^{-2}$ , and 4200 s for along-grain ignition at  $12 \text{ kW m}^{-2}$ . On this basis, one might conclude that  $7.5 \text{ kW m}^{-2}$  is  $\dot{q}''_{\min}$  for piloted ignition of wood, but the value for piloted ignition cannot be higher than for autoignition and the latter may be low indeed (see below).

Generally, ‘piloted ignition’ means the presence of a flame or a spark in the gas phase where pyrolysates accumulate. But it is also possible to apply a gas flame directly onto a surface as an ‘impinging pilot,’ in which case much less radiant heating is needed to achieve ignition since a local heat flux concentration is created. An old Fire Research Station study [65] showed  $\dot{q}''_{\min} = 5 \text{ kW m}^{-2}$  for Western red cedar and Douglas fir. No other published studies exist. Apart from surface-applied pilots, both the type of pilot and its location can affect ignition times. Several studies [57,66] produced limited



data – more studies would be needed to quantify trends reliably. It is also possible to heat a wood surface by applying a relatively-uniform ‘wall of flame’ onto it, and this is discussed later.

### Experimental Results on Autoignition

Unlike piloted ignition, autoignition of wood under radiant heating conditions has been studied by only a few researchers, most notably Simms, who conducted various experiments at FRS in the 1950s and ’60s. In a 1952 study, he tested 6 different species of wood using 19 mm thick specimens [67]. The results, including the correction for a 20% flux mis-calibration [68], are shown in Table 4. In a 1961 study [69], he reported an enormous value of up to  $117 \text{ kW m}^{-2}$  for autoignition of blackened oak and cedar specimens. In a 1967 study [70], he reported minimum fluxes for piloted ignition that were similar to the corrected 1952 values, but autoignition values reported were quite a bit higher, being approximately  $40\text{--}50 \text{ kW m}^{-2}$ . In his 1961 study, Simms noted that a draft strong enough to be turbulent was helpful in reducing the  $\dot{q}''_{\min}$ . This was evidently a gas-phase effect, but even today there is no systematic knowledge on gas-phase ignition effects. In another study [71], Simms concluded that the quantitative effect of the

**Table 4. Minimum flux for autoignition of wood, as reported by various researchers.**

Study	Orient.	MC (%)	Draft	Specimen Size Exposed	Max. Time of Test	$\dot{q}''_{\min}$ ( $\text{kW m}^{-2}$ )	Notes
Lawson, Simms [67]	V	0	N	50 × 50 mm	20 min	29–33	
Simms [69]	V	0	Y N	8 mm $\phi$	14 s 18 s	75–100 117	black-ened surface
Simms, Law [70]	V	0	N	76 × 76 mm 150 × 150 mm	70 s 79 s	46 42	
Moran [15]	V	0	Y	50 × 50 mm	9 min	25	
Shields et al. [57]	H	≈ 10	N	100 × 100 mm	96 s	30–40	
	V				59 s	40–50	
	H	≈ 10	N	165 × 165 mm	12 min	< 20	ISO 5657 test
Shoub, Bender [18]	V	≈ 10	N	920 × 920 mm	3.9–5.2 h	4.3	

rather small exposure size of 8 mm is nearly negligible, so presumably the enormous  $\dot{q}''_{\min}$  values in the 1961 study were mainly due to insufficiently long test time.

Moran [15] examined the ignition of vertical panels of 6.4 mm thick ponderosa pine using an electric radiant panel and found  $\dot{q}''_{\min} = 25 \text{ kW m}^{-2}$ . Shields et al. [57] examined the autoignition of Sitka spruce in the Cone Calorimeter and in the ISO 5657 apparatus. They exposed specimens in increments of  $10 \text{ kW m}^{-2}$ , so their results were only approximate. Since the heater arrangements have some similarity, it is not clear why the values obtained in the Cone Calorimeter and the ISO 5657 apparatus were not closer. Shields' data does illustrate that it is much more difficult to achieve autoignition in the vertical orientation than in the horizontal orientation. The above studies were all of less than 20 min duration. Only the study by Shoub and Bender [18] involved longer-term exposures. They used an electric radiant panel operating at an effective black-body face temperature of  $273^\circ\text{C}$  and producing a heat flux of  $4.3 \text{ kW m}^{-2}$  at the center of the specimen, and lower heat fluxes at the edges. While they did not test any whole woods, they tested 13 mm plywood. It ignited at the  $4.3 \text{ kW m}^{-2}$  flux, but required waiting over 5 h. In their tests, they also documented that the face temperatures of the specimens in some cases reached temperatures higher than that of the radiant source, indicating that self-heating of the material was important and that assuming an inert solid would not be appropriate. It should be of high priority that modern-day researchers attempt to repeat these experiments and verify their results. The conclusion – pending a verification of Shoub and Bender's [18] results – is that wood will autoignite at about  $4.3 \text{ kW m}^{-2}$ , if exposed for hours, rather than minutes. For short-term exposures, a value of  $20 \text{ kW m}^{-2}$  perhaps best captures the research results.

At any given irradiance, if ignition occurs under both autoignition and piloted ignition conditions, it is evident that ignition times for the latter will be shorter (unless the pilot is badly placed). A tractable theory, such as Janssens', models only the solid phase, so the presumed conclusion would be that ignition times do not change. A more refined point of view would be to assume that for autoignition, heating up the solid to the same temperature suffices as for the piloted case, but that afterwards a delay time must be added to account for gas-phase events. A theory of this sort has not been developed, however. Experimentally, even though there is a great deal of scatter (Figure 5), the results of Shields et al. [57] can be used to derive an equation:

$$t_{ig}(\text{autoignition}) = (2.86 - 0.0172\dot{q}''_e) \cdot t_{ig}(\text{spark})$$

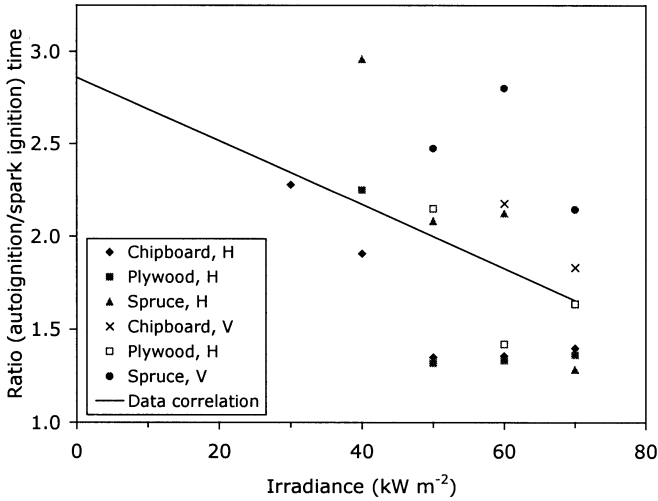


Figure 5. Ratio of ignition times, autoignition/sparkignition.

Thus, for example, at a flux of  $25 \text{ kW m}^{-2}$ , under autoignition, ignition times can be expected to be 2.43 times those for the spark-ignition case, while at  $50 \text{ kW m}^{-2}$  the factor drops down to 2 times. Since the highest experimental flux was  $70 \text{ kW m}^{-2}$ , the rule should not be extrapolated to greatly higher fluxes. Also, due to the data scatter, the guidance is only semiquantitative.

### IGNITION FROM MISCELLANEOUS HEAT SOURCES

There is very little data on ignition of wood from flames, despite the fact that this is how we light our fireplaces. When a thin piece of wood is lit at the bottom, burning may continue to completion. But a thick piece of wood will not undergo self-sustained combustion under the same circumstances. Bryan [72] reports that the maximum thickness for self-sustained burning, given a flaming ignition at the bottom of a vertical piece, is about 19 mm. In a horizontal orientation, even 12 mm thick specimens have been found to be too thick for self-sustained burning [73]. Using the methenamine pill test (a standard test for floor coverings), it was found [74] that no ignition occurs for any of a wide variety of wood products tested in thicknesses of 10–21 mm. Ignitability of wood boards has also been examined [75] using the ISO 11925-2 small-burner test. Using a 30 s flame exposure to the surface, ignition rarely occurred and never spread to the 150 mm limits, even with specimens as thin as 2 mm. For 30 s bottom-edge impingement, specimens of 18 mm thickness or less commonly ignited, but only ones of 10 mm thickness or less generally reached the 150 mm mark.

Ebeling [76] studied the ignition of wood panels when exposed to a flame, with the flame being applied against the whole face. He tested oak, white pine, redwood, and yellow pine, and his results can be represented by the correlation:

$$t_{ig} = 41.3\rho^{0.94}(\dot{q}''_e)^{-1.82}$$

although there was a wide spread of results. The above equation implies that the critical flux is identically zero, which is at least partly due to the fact that there is no convective cooling of the surface in a flame-ignition test. For the same reason, their flame ignition times were a fair bit shorter than times obtained by applying the same heat flux in a radiant heating test.

Even though convective heating is an important feature in a Setchkin-type apparatus, there has been no scientific study where ignition would be primarily from convective heating.

When a sufficiently high voltage is impressed across a tree or a wood member, an arc tracking process takes place. Wood first dries out at the electrodes, then a carbonized channel starts to form. Given enough time and voltage, sufficient heating of the carbonized track takes place that the electric current passing through the track heats up the wood to ignition. This process has been studied by several researchers [77–80]. It is mainly of concern in connection with high voltage wiring, including power line poles and neon signs installed near wood surfaces.

Ignition with laser radiation produces very different results than radiation from flames or grey-body radiators, for reasons not fully explained. Kashiwagi [81] exposed horizontally-oriented red oak specimens to laser radiation at  $10.6\mu\text{m}$  and found high values of  $\dot{q}''_{\min} \approx 80\text{ kW m}^{-2}$  for autoignition and  $55\text{ kW m}^{-2}$  for piloted ignition. This is also common with laser ignition of other substances, e.g., plastics. Ignition from nuclear weapons has been simulated [82] by use of radiant exposure from an arc-image furnace. The results for the brief, high-intensity pulses were expressed in terms of energy fluence. Using 13 mm thick Douglas fir, transient flaming was observed for an energy fluence of  $1090\text{ kJ m}^{-2}$  (480 kt bomb) and sustained flaming at  $1300\text{ kJ m}^{-2}$  (1180 kt) or higher. Yellow poplar of 1.6 mm thickness also showed sustained flaming at  $1090\text{ kJ m}^{-2}$ , but no transient flaming regime.

## IGNITION FROM HOT BODIES, FIREBRANDS, AND SMOLDERING

Glowing and smoldering are similar, but not identical mechanisms of ignition. Smoldering is, by definition, a self-sustained process. Ignition and consumption of a wood material by glowing, on the other hand, can occur if

it is subject to sufficient radiant or convective heating, without a requirement that the process continue, should the external heat source be removed. Firebrands themselves may be flaming or glowing, and they may, in some cases, initially cause flaming in the target fuel, although a smoldering ignition is the usual concern.

Self-sustained smoldering occurs easily in various wood products which are highly porous or finely divided (fiberboard, wood shavings, rotted wood, etc.). Whole wood, however, is only slightly porous to the inflow of oxygen and will not smolder as a single surface facing open air. Ohlemiller [83] reports that by supplying external heating at approximately  $10 \text{ kW m}^{-2}$ , wood can be made to burn in a glowing mode; this of course is not smoldering, since it is not self-sustained. By preheating the bulk of the wood sufficiently, continued combustion can be maintained. This can be seen in a fireplace where individual pieces may continue glowing even after a 'three-log' effect no longer exists. Only a limited number of experimental studies exist on the question of minimum conditions necessary to start wood smoldering. Ohlemiller [84] conducted experiments where smoldering was achieved by providing a 'three-log' arrangement and igniting the surfaces with flat electric heaters. Even with the optimal geometry, air flow velocity had to be within a close range for sustained smoldering to be seen.

Solid wood is most commonly ignited by firebrands during wildland fires. Humidity plays a strong role in the process, and wildland fires often involve extremes of high temperature, low humidity, and strong wind gusts. Only a few laboratory studies have been conducted on the ignitability of solid wood by firebrands. CSIRO researchers [85,86] found that some surprisingly small (0.8–12 g) wood cribs sufficed for ignition. An inside-corner ('re-entrant corner') geometry of the siding was especially conducive to ignition. Hamada et al. [87] found that under no-wind conditions, red-hot brands of about 5 mm diameter caused ignition, but in an  $8 \text{ m s}^{-1}$  wind, even brands of 2.5 mm were likely to cause ignition. Low RH values (20%) were needed for this to occur. Applying flames to the surface of a wood structural member will not result in smoldering ignition, unless the flame is applied for so long that the wood member is largely burned up. Specifically, it has been demonstrated [88] that applying the flame from an acetylene/air plumber's torch directly onto wood studs for periods of 1–5 min leads to local charring but no sustained combustion of any type once the torch is removed and the flames self-extinguish.

A special problem is one where ignition of wood occurs from steam pipes or from a metal heating system part which is in contact with the wood for a long time. Under long-term heating (months-to-years), it appears that wood can ignite when a surface is held at a temperature lower than the ignition temperature determined from tests that last a short time (minutes-to-days).

The information largely comes from case histories and good experiments are lacking. Babrauskas [89] recently reviewed the state of the art on this topic.

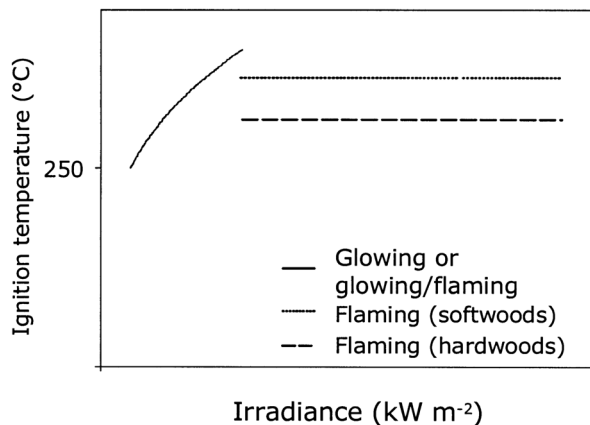
## CONCLUSIONS

Some aspects of the wood ignition problem are well-known, and these can be used in routine engineering applications. This is primarily true of ignition times for piloted ignition, provided that fluxes too close to the minimum are not considered. But despite more than a century of scientific research, many other aspects of wood ignition are poorly known. Ignition temperature values published in the literature encompass a very large range, but the present study has identified that this is mostly because the ignition temperature is a function of the heat flux and the mode of ignition, and is not a universal constant. Concerning the two innately different ways of defining ignition temperature – as the temperature of the specimen, versus the furnace, at the moment of ignition – it appears that, for wood, this difference is generally small enough to be obscured by experimental data scatter.

Conclusive data now exist to conclude that at the minimum flux and for a certain range of heat fluxes above that, ignition of wood occurs as a glowing ignition. For higher fluxes, ignition is a two-stage process, with initial glowing ignition later transitioning to flaming. At still higher fluxes, wood ignites directly in the flaming mode. Despite this qualitative knowledge, the pertinent flux ranges are not yet well quantified. Analysis of available data leads to the summary given in Table 5 and schematically depicted in Figure 6. High fluxes (e.g., over  $80 \text{ kW m}^{-2}$ ) are not listed since there is little information, but also this regime is of less interest to fire safety. Most researchers have conducted much too short tests in attempting to define ‘minimum’ conditions, thus only a single study is the basis for observing that ignition may occur at heat fluxes as low as  $4.3 \text{ kW m}^{-2}$ . It is, of course, likely that there is not a unique minimum flux value, but that various factors – apart from inadequate duration of experiments – can affect its value. Likewise, the flux value (perhaps  $22 \text{ kW m}^{-2}$  for piloted,  $33 \text{ kW m}^{-2}$  for

**Table 5. Summary of ignition temperature results.**

Flux Ignition type	Minimum glowing or 2-step	Low	Medium Flaming
$T_{ig}$ (°C), piloted	250	350–400 peak, lower for fluxes close to minimum.	300–310 hardwoods 350–365 softwoods
$T_{ig}$ (°C), autoignition	250	no data	380–400 ??



**Figure 6.** The effect of irradiance of piloted ignition temperature.

autoignition) dividing the medium flux (flaming ignition) from low flux (ignition starts with glowing) regimes needs better study.

It is also noteworthy that there exists no theoretical study on glowing ignition that is directly applicable to wood and only limited studies of other substances. Somewhat related to the lack of knowledge about glowing ignitions is the lack of knowledge on ignitions from hot bodies. Experimental data on this topic are so scarce that it can only be concluded that ignitions are possible under some surprisingly mild attacks, e.g., firebrands of a few grams. Ignition of wood in actual fires often is due to direct flame contact with the material but, again, guidance on this topic is minimal.

## REFERENCES

1. "Standard Test Method for Determining Ignition Temperature of Plastics (ASTM D 1929)," American Society for Testing and Materials, West Conshohocken, PA.
2. Hill, H.B. "On the Behavior of Sound and Decayed Wood at High Temperatures," Proc. Amer. Acad. Arts and Sciences, Vol. 22, 1887, pp. 482–492.
3. Bixel Edward, C. and Moore, Howard, J. "Are Fires Caused by Steam Pipes?," (B.S. Thesis), Case School of Applied Science, Pittsburgh, 1910.
4. Banfield, W.O. and Peck, W.S. "The Effect of Chemicals on the Ignition Temperature of Wood," Canadian Chemistry and Metallurgy, Vol. 6, 1922, pp. 172–176.
5. Brown, C.R. "The Determination of the Ignition Temperatures of Solid Materials," (D.Sc. Thesis), The Catholic University of America, Washington, 1934.
6. VanKleeck, A. "A Preliminary Study of Ignition Temperatures of Finely Chopped Wood (Project L-179)," Forest Products Lab., Madison WI, 1936.
7. As cited in: "Wood and Wood-Based Products," J.M. Cholin, ed., Fire Protection Handbook, A.E. Cote and J.L. Linville, eds., 18th Edn., National Fire Protection Assn., Quincy MA, 1997.

8. Graf, S.H. "Ignition Temperatures of Various Papers, Woods, and Fabrics," Oregon State College Bull, 26, Oregon State College, Corvallis, March 1949.
9. Angell, H.W., Gottschalk, F.W. and McFarland, W.A. "Ignition Temperature of Fireproofed Wood, Untreated Sound Wood and Untreated Decayed Wood," British Columbia Lumberman, Vol. 33, 1949, pp. 57-58, 70-72.
10. Fons, W.L. "Heating and Ignition of Small Wood Cylinders," Ind. and Eng. Chem., Vol. 42, 1942, pp. 2130-2133.
11. Narayanamurti, D. "A Note on Pyrolysis and Ignition of Wood," Current Science, Vol. 27, 1956, pp. 22-23.
12. Thomas, P.H., Simms, D.L. and Theobald, C.R. "The Interpretation of Some Experimental Data on the Ignition of Wood," Fire Research Note No. 411, Fire Research Station, Borehamwood, UK, 1959.
13. Akita, K. Studies on the Mechanism of Ignition of Wood," Report of the Fire Research Institute of Japan, Vol. 9, 1959, pp. 1-44, 51-54, 77-83, 99-105.
14. Simms, D.L. "Ignition of Cellulosic Materials by Radiation," Combustion and Flame, Vol. 4, 1960, pp. 293-300.
15. Moran, H.E. jr. "Effectiveness of Water Mists for Protection from Radiant Heat Ignition," NRL Report 5439, US Naval Research Laboratory, Washington, 1960.
16. Patten, G.A. "Ignition Temperatures of Plastics," Modern Plastics, Vol. 38, July 1961, pp. 119-122, 180.
17. Buschman, A.J. "Ignition of Some Woods Exposed to Low Level Thermal Radiation," NBS Report 7306, Natl. Bur. Stand., Washington, DC, 1961.
18. Shoub, H. and Bender, E.W. "Radiant Ignition of Wall Finish Materials in a Small Home," NBS 8172, Natl. Bur. Stand, Washington, DC, 1964.
19. Tinney, E.R. "The Combustion of Wooden Dowels in Heated Air," in 10th Symp. (Intl.) on Combustion, The Combustion Institute, Pittsburgh, 1964, pp. 925-930.
20. Simms, D.L. and Law, M. "The Ignition of Wet and Dry Wood by Radiation," Combustion and Flame, Vol. 11, 1967, pp. 377-388.
21. Muir, W.E. "Studies of Fire Spread between Buildings," (PhD Dissertation), University of Saskatchewan, Saskatoon, Canada, 1967.
22. Koohyar, A.N. "Ignition of Wood by Flame Radiation," (PhD Dissertation), Univ. Oklahoma, Norman, 1967.
23. Melinek, S.J. "Ignition Behaviour of Heated Wood Surfaces," FR Note 755, Fire Research Station, Borehamwood, UK, 1969.
24. Jach, W. "Das Verhalten von Holz und Holzwerkstoffen bei Dauereinwirkung von Temperaturen unterhalb des Flamm- und Brennpunktes," Mitteilungen der deutschen Gesellschaft für Holzforschung, Nr. 56, 12-17, 1969.
25. Smith, W.K. and King, J.B. "Surface Temperatures of Materials during Radiant Heating to Ignition," J. Fire and Flammability, Vol. 1, 1970, pp. 272-288.
26. Atreya, A. "Pyrolysis, Ignition and Fire Spread on Horizontal Surfaces of Wood," (PhD Dissertation), Harvard University, Cambridge, MA, 1983.
27. Atreya, A., Carpentier, C. and Harkleroad, M. "Effect of Sample Orientation on Piloted Ignition and Flame Spread," in Fire Safety Science - Proc. 1st Intl. Symp., Hemisphere, Washington, 1986, pp. 97-109.
28. Abu-Zaid, M. "Effect of Water on Ignition of Cellulosic Materials," (Ph.D. Thesis), Michigan State Univ., East Lansing, MI, 1988.
29. Janssens, M.L. "Fundamental Thermophysical Characteristics of Wood and Their Role in Enclosure Fire Growth," (PhD dissertation), University of Gent, Belgium, 1991.
30. Li, Y. and Drysdale, D. "Measurement of the Ignition Temperature of Wood," in Fire Science and Technology - Proc. First Asian Conf., Intl. Academic Publishers, Beijing, 1992, pp. 380-385.



31. Masařík, I. "Ignitability and Burning of Plastic Materials: Testing and Research," in *Interflam '93*, Interscience Communications Ltd., London, 1993, pp. 567–577.
32. Fangrat, J., Hasemi, Y., Yoshida, M. and Hirata, T. "Surface Temperature at Ignition of Wooden Based Slabs," *Fire Safety J.*, Vol. 27, 1996, pp. 249–259; Vol. 28, 1997, pp. 379–380.
33. Moghtaderi, B., Novozhilov, V., Fletcher, D.F. and Kent, J.H. "A New Correlation for Bench-Scale Piloted Ignition Data of Wood," *Fire Safety J.*, Vol. 29, 1997, pp. 41–59.
34. "Standard Test Method for Heat and Visible Smoke Release Rates for Materials and Products using an Oxygen Consumption Calorimeter (E 1354)," American Society for Testing and Materials, West Conshohocken PA.
35. Spearpoint, M.J. "Predicting the Ignition and Burning Rate of Wood in the Cone Calorimeter Using an Integral Model," NIST GCR 99–775, Nat. Inst. Stand. and Technol., Gaithersburg MD, 1999.
36. Prince, R.E. "Tests on the Inflammability of Untreated Wood and of Wood Treated with Fire-Retarding Compounds," *Proc. NFPA*, Vol. 19, 1915, pp. 108–158.
37. Buchanan, M.A. "The Ignition Temperature of Certain Pulp and Other Wood Components," *TAPPI*, Vol. 35, 1952, pp. 209–211.
38. Urbas, J. and Parker, W.J. "Surface Temperature Measurements on Burning Wood Specimens in the Cone Calorimeter and the Effect of Grain Orientation," *Fire and Materials*, Vol. 17, 1993, pp. 205–208.
39. Shafizadeh, F. "Utilization of Biomass by Pyrolytic Methods," in *Proc. TAPPI 1997 Joint Forest Biology/Wood Chemistry Mtg.*, Madison WI, 1977, pp. 191–199.
40. Bilbao, R. et al. "Experimental and Theoretical Study of the Ignition and Smoldering of Wood Including Convective Effects," *Combustion and Flame*, Vol. 126, 2001, pp. 1363–1372.
41. Janssens, M. "Piloted Ignition of Wood: A Review," *Fire and Materials*, Vol. 15, 1991, pp. 151–167.
42. Baer, A.D. and Ryan, N.W. "Ignition of Composite Propellants by Low Radiant Fluxes," *AIAA J.*, Vol. 3, 1965, pp. 884–889.
43. Lengellé, G., Bizot, A., Duterque, J. and Amiot, J.-C. "Ignition of Solid Propellants," *La Recherche Aérospatiale – English Edition*, No.2, 1991, pp. 1–20.
44. Moussa, N.A., Toong, T.Y. and Garris, C.A. "Mechanism of Smoldering of Cellulosic Materials," in 16th Symp. (Intl.) on Combustion, The Combustion Institute, Pittsburgh, 1976, pp. 1447–1457.
45. Fredlund, B. "A Model for Heat and Mass Transfer in Timber Structures during Fire," Report LUTVDG/TVGG-1003, Dept. of Fire Safety Engineering, Lund University, Lund, Sweden, 1988.
46. Ohlemiller, T.J. "Modeling of Smoldering Combustion Propagation," *Prog. Energy and Comb. Sci.*, Vol. 11, 1985, pp. 277–310.
47. Khitrin, L.N. "Physics of Combustion and Explosion," Israel Program for Scientific Translations, Jerusalem, 1962.
48. Smith, I.W. "The Combustion Rates of Coal Chars: A Review," in 19th Symp. (Intl.) on Combustion, The Combustion Institute, Pittsburgh, 1982, pp. 1045–1065.
49. Yuen, R.K.-K. "Pyrolysis and Combustion of Wood in a Cone Calorimeter," (PhD Dissertation), The University of New South Wales, Australia, 1998.
50. Long, R.T. jr., Torero, J.L., Quintiere, J.G. and Fernandez-Pello, A.C. "Scale and Transport Considerations on Piloted Ignition of PMMA," in *Fire Safety Science – Proc. 6th Intl. Symp.*, Intl. Assn. for Fire Safety Science, 2000, pp. 567–578.
51. Hu, X.-F. and Clark, F.R.S. "The Use of the ISO/TC92 Test for Ignitability Assessment," *Fire and Materials*, Vol. 12, 1988, pp. 1–5.
52. Mikkola, E. "Ignitability Comparisons between the ISO Ignitability Test and the Cone Calorimeter," *J. Fire Sciences*, Vol. 9, 1991, pp. 276–284.

53. Nussbaum, R.M. and Östman, B.A.-L. "Larger Specimens for Determining Rate of Heat Release in the Cone Calorimeter," *Fire and Materials*, Vol. 10, 1986, pp. 151–160.
54. Hallman, J.R. "Ignition Characteristics of Plastics and Rubber," (PhD Dissertation), University of Oklahoma, Norman, 1971.
55. Peck, E.C. "Moisture Content of Wood in Dwellings," Circular No. 239, US Dept. of Agriculture, Washington, 1932.
56. Henderson, A. "Predicting Ignition Time Under Transient Heat Flux Using Results from Constant Heat Flux Experiments," Fire Engineering Research Report 98/4, School of Engineering, University of Canterbury, New Zealand, 1999.
57. Shields, T.J., Silcock, G.W. and Murray, J.J. "The Effects of Geometry and Ignition Mode on Ignition Times Obtained Using a Cone Calorimeter," *Fire and Materials*, Vol. 17, 1993, pp. 25–32.
58. Mikkola, E. "Puupinnan syttyminen," Research Notes 1087, Valtion Teknillinen Tutkimuskeskus, Espoo, Finland, 1989.
59. Dlugogorski, B., Pope, D.M., Moghtaderi, B., Kennedy, E.M. and Lucas, J.A. "A Study on Fire Properties of Australian Eucalyptus," in *Wood and Fire Safety – 4th Intl. Scientific Conf.*, The High Tatras, Slovak Republic, 2000, pp. 57–72.
60. Janssens, M.L. private communication, 2001.
61. Babrauskas, V. unpublished NIST test results.
62. Grexa, O. private communication, 2001.
63. Wesson, H.R. "The Piloted Ignition of Wood by Radiant Heat," (D.Eng. dissertation), Univ. Oklahoma, Norman, 1970.
64. McGuire, J.H. *Fire and the Spatial Separation of Buildings*, Fire Technology, Vol. 1, 1965, pp. 278–287.
65. Pickard, R.W., Simms, D.L. and Walters, J.E.L. "The Ignition by Radiation of Wood Protected by Some Common Paints," FPE Note No. 61/1951, Fire Research Station, Borehamwood, 1951.
66. Simms, D.L. and Hird, D. "On the Pilot Ignition of Materials by Radiation," FR Note 365, Fire Research Station, Borehamwood, England, 1958.
67. Lawson, D.I. and Simms, D.L. "The Ignition of Wood by Radiation," *Brit. J. Appl. Phys.*, Vol. 3, 1952, pp. 288–292.
68. Simms, D.L. "On the Pilot Ignition of Wood By Radiation," *Combustion and Flame*, Vol. 7, 1963, pp. 253–261.
69. Simms, D.L. "Experiments on the Ignition of Cellulosic Materials by Thermal Radiation," *Combustion and Flame*, Vol. 5, 1961, pp. 369–375.
70. Simms, D.L. and Law, M. "The Ignition of Wet and Dry Wood by Radiation," *Combustion and Flame*, Vol. 11, 1967, pp. 377–388.
71. Simms, D.L. "Ignition of Cellulosic Materials by Radiation," *Combustion and Flame*, Vol. 4, 1960, pp. 293–300.
72. Bryan, J. "The Fire Hazard," *Wood*, Vol. 8, 1943, pp. 260–262.
73. Wilson, J.A. "A Different View on Plastics Fire Hazard Classification," *NFPA Q.*, Vol. 55, October 1962, pp. 162–164.
74. Tsantaridis, L., "CEN Tablet Test Results for Wood Floorings," L-Rapport 9702009, Trätec, Stockholm, 1997.
75. Tsantaridis, L., "CEN Ignitability Test Results for Wood Building Products," L-Rapport 9702010, Trätec, Stockholm, 1997.
76. Ebeling, K.L. "Ignition of Wood by Direct Flame Contact," (M.S. Thesis), Univ. Oklahoma, Norman, 1973.

77. Kinbara, T. and Sue, S. "On the Outbreak of Fire Due to Leakage of Electricity from Neon Transformer through Planking," [in Japanese], Bull. Fire Prev. Soc. of Japan, Vol. 2, No. 2, 1953, pp. 39–41.
78. Kinbara, T. and Takizawa, K. "Ignition of a Salt-soaked Wooden Board by an Electric Current through It," [in Japanese], Bull. Fire Prev. Soc. of Japan, Vol. 11, No. 2, Dec. 1961, pp. 26–31.
79. Sanderson, J.L. "Carbon Tracking: Poor Insulation Combined with Contaminants is Potential Fire Cause," Fire Findings, Vol. 8, No. 3, 2000, pp. 1–3.
80. Blackburn, T.R. and Pau, L.T. "Characteristics of a Simulated High Impedance Fault," in Electric Energy Conf. 1985, Newcastle, Australia, 1985, pp. 301–304.
81. Kashiwagi, T. "Experimental Observation of Radiative Ignition Mechanisms," Combustion and Flame, Vol. 34, 1979, pp. 231–244.
82. "Critical Radiant Exposures for Ignition of Tinder and Combustible Materials (Part I – Wood)," Naval Applied Science Lab., Brooklyn NY, 1965.
83. Ohlemiller, T.J. "Smoldering Combustion," NBSIR 85–3294, Natl. Bur. Stand., Gaithersburg, MD, 1986.
84. Ohlemiller, T.J. "Smoldering Combustion Propagation on Solid Wood," in Fire Safety Science – Proc. 3rd Intl. Symp., Elsevier, New York, 1991, pp. 565–574.
85. McArthur, N.A. and Lutton, P. "Ignition of Exterior Building Details in Bushfires: An Experimental Study," Fire and Materials, Vol. 15, 1991, pp. 59–64.
86. Dowling, V.P. "Ignition of Timber Bridges in Bushfires," Fire Safety J., Vol. 22, 1994, pp. 145–168.
87. Hamada, M. et al. "Experiments on the Ignition due to Fire Brands," [in Japanese], Fire Research – Reports from the Fire Science Research Committee, Property and Casualty Insurance Rating Organization of Japan, Tokyo, 1951.
88. Babrauskas, V. and Hall, J. unpublished test work, 2000.
89. Babrauskas, V. "Pyrophoric Carbon. . .The Jury is Still Out," Fire and Arson Investigator, Vol. 51, No. 2, Jan. 2001, pp. 12–14.